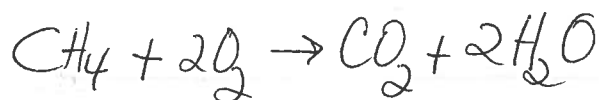


Honors Chemistry

Daily Quiz

Dec 8/9, 2009

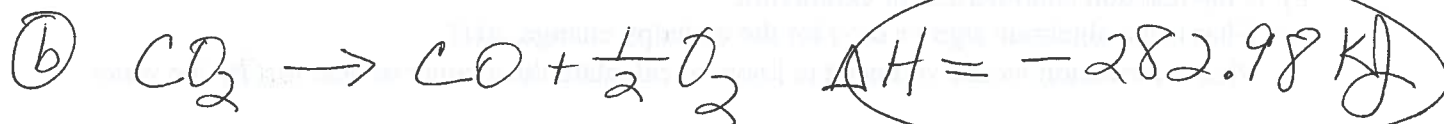
From the text:



$$(52) \quad q = mc\Delta T \quad q_{\text{H}_2\text{O}} = (250\text{g})(4.18)(25^\circ\text{C}) = 26,125\text{J}$$

$$26,125\text{kJ} \times \frac{1\text{mol CH}_4}{-802.3\text{kJ}} = 3.26 \times 10^{-2}\text{mol CH}_4$$

$$3.26 \times 10^{-2}\text{mol CH}_4 \times \frac{16.05\text{g}}{1\text{mol CH}_4} = \boxed{0.52\text{g CH}_4}$$



✓ Prof Crum